reactions was determined at widely different temperatures. At high temperatures, the rates of the two reactions are approximately the same and their energies of activation are equal. This finding indicates that at high temperatures the limiting step is identical for both reactions.

Calcination of the catalyst in vacuum at a temperature of 700° imparts to it high catalytic activity —even at −78°— in the homomolecular exchange; however, subsequent heating of the catalyst in oxygen at temperatures of over 200° completely deactivates the catalyst. At low temperatures, the exchange reactions occur in two temperature regions, which are characterized by different activation energies.

Effect of Surface Acidity of Catalysts on Rate of Polymerization of Three-Membered Heterocyclics

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Polymerization of ethylene oxide, propylene oxide, and ethylenimene in gaseous phase was determined gravimetrically over several Al_2O_3— and $SiO_2 \cdot Al_2O_3$—based heterogeneous catalysts at temperatures of 0°–95°. Pretreatment of the Al_2O_3 base with HCl and BF_3 decreases activity of the Al_2O_3-based catalysts in polymerization of the olefin oxides and increases it in polymerization of ethylenimene. Silica-alumina is less active catalytically than aluminum oxide in polymerization of the olefin oxides. Calcination of this support at 800°–900° increases its activity in these reactions. Replacement of the H^+ ions on the silica-alumina surface with Mg^{+2}, Ca^{+2}, and Sr^{+2} increases activity of this support in polymerization of the olefin oxides but decreases it in polymerization of ethylenimine. The results obtained in the catalytic polymerization of the olefin oxides are attributed to progress of the reaction by the coordination mechanism and those in the polymerization of ethylimine, to progress by the proton-acidic mechanism.

Mechanism and Kinetics of Catalytic Conversion of Acetylene: Kinetics of Acetonitrile Synthesis from Acetylene and Ammonia Over a ZnO/Al_2O_3 Catalyst

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A study of kinetics of acetonitrile synthesis from acetylene and ammonia was carried out in a nongradient reaction system, using a zinc oxide catalyst at temperatures of 340°–435°. The kinetic equation based on the experimental data satisfactorily describes the reaction at the conditions studied. Analysis of the experimental data and of the kinetic equation validates the assumption that the limiting step of the process is dehydrogenation of the adsorbed intermediate product, ethylidenimine, which is formed by the interaction of acetylene with ammonia.

Kinetics of Cyclohexane Splitting by Steam Over Nickel Catalysts

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Activity of a Ni/C and a Ni/SiO_2 catalysts was investigated in cyclohexane-water reaction at temperatures of 400°–460°. Despite the near-equal activity of the two catalysts in dehydrogenation of cyclohexane, the Ni/C catalyst—unlike to the Ni/SiO_2 catalyst—is totally inactive in the presence of water. Over the Ni/SiO_2 catalyst, the reaction is of the zero order with respect to the cyclohexane; with respect to the water, the order of the reaction changes with increasing water concentration from that of the first to that of the zero. Activation energy of the cyclohexane-water reaction over the Ni/SiO_2 catalyst is 22–24 kcal/mol.

Kinetics of Copolymerization Reactions in Continuous Flow Systems at Conditions of Ideal Mixing

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The kinetics of copolymerization reactions in continuous flow systems were analyzed at the conditions of ideal mixing in the reactor. The equations derived on the basis of the data of the batch-type studies are adequate to calculate conversion, composition of the copolymer, and the degree of polymerization at each specified set of conditions. These equations also aid in determining the parameters of the processing variables to obtain the copolymers with properties desired. The sample calculations are shown for copolymerization of styrene with methylmethacrylate in the presence of azoisobutyric acid at temperatures of 60°–80°.